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The Synthesis of Alum Lab Michaela
Tonsager and Kaili Johnson
Conclusion We determined that our
sample was in fact alum. Our melting
point of 99.4 degrees C was similar to
the published melting point of 92.5
degrees C. Our percent sulfate was

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42.44%, which is close to the The
Synthesis of Alum Lab by Michaela
Tonsager

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The aluminum is being recycled in the sense that the aluminum undergoes a process that adapts it to a new function, instead of just converting the shape of the metal. The aluminum in this experiment is converted to alum $[KAl(SO_4)_2 \cdot 12H_2O]$ which

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is the usual term for a type of compound with the general formula $M_2(SO_4)_3 \cdot 12H_2O$. M is a monovalent cation, and M' is a trivalent cation, in this case, Al^{3+} .

~~Recycling Aluminum lab write up:
experiment 3 - CHEM 2070 ...~~

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Lab 4 synthesis and analysis of alum
part. Alum from the synthesis of
potassium aluminum sulfate chem lab

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page for the Rinse the reaction beaker twice with 2 mL distilled water, and add each rinse to the filter paper. Allow water to run through the aspirator to establish a vacuum in the filter flask. Red numbers are variable.

~~Synthesis of alum preliminary lab~~

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~~assignment answers ...~~

1. Either bring your own aluminum can or use the pieces provided in the lab. You will need a piece of scrap aluminum about 7.5 cm by 5.0 cm that weighs 1.0 to 1.1 grams. 2. Scrape both sides of the piece (using sandpaper) to remove the paint and

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lacquer (this will speed the dissolving process GREATLY, and give a more accurate starting mass). 3.

~~Synthesis of Alum from Aluminum~~
Synthesis of Alum Alum is a solid ionic compound with many uses. It is used as astringent to prevent bleeding from

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small cuts, as an ingredient in deodorants, as an ingredient in baking powders, and as a preservative used in pickling. The formula of alum is $KAl(SO_4)_2 \cdot 12H_2O$ and the standard name is potassium aluminum sulfate. It is an

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~~CHEM 231 Experiment 5 Synthesis of
Alum~~

Numbers of moles of Alum = 0.0397
moles Molar mass of Alum = 474.39 g
/ mole Mass of Alum yielded =
Number of moles * Molar mass of
Alum 4. Mass of Alum yield = 0.0397
moles * 474.39 (g / mole) = 18.83 g %

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yield = (Actual yield (g) / Theoretical yield (g)) * 100 = (14.14 g / 18.83 g) * 100 = 75.1 % Discussion: The experimental value for the yielding of alum is 14.14 g and theoretical value is 18.83 g.

~~Lab report on synthesis of Alum using~~

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Aluminum:

In this experiment you will prepare and characterize alum (potassium aluminum sulfate dodecahydrate, $KAl(SO_4)_2 \cdot 12 H_2O$). The first step in this synthesis, which you will perform during Week 1, is to react metallic aluminum with a

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concentrated solution of potassium hydroxide (KOH) to form the potassium salt of the tetrahydroxoaluminate complex ion, $[\text{Al}(\text{OH})_4]^-$.

~~Preparation and Analysis of Alum |
Chem Lab~~

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CHEM 231 Experiment 5 Synthesis of Alum Alum is aluminum sulfate.

Aluminum reacts with sulfuric acid to produce alum. Chem Lab: Synthesis of Alum Question!? | Yahoo Answers The overall reaction to form alum is: $\text{Al}^{3+}(\text{aq}) + \text{K}^{+}(\text{aq}) + 2 \text{SO}_4^{2-}(\text{aq}) + 12 \text{H}_2\text{O} \rightarrow \text{KAl}(\text{SO}_4)_2 \cdot 12\text{H}_2\text{O}(\text{s})$

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Crystallization. Using tongs, remove the filtrate beaker to a wire

~~Alum Synthesis Lab Answers~~
~~catalog.drapp.com.ar~~

1. Aluminum is the third most abundant element in the earth's crust.
2. The supply of aluminum ore is not

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inexhaustible. 3.The winning, or extraction of the metallic form an impure ionic source, of aluminum from aluminum ore is very costly from an energy point of view. 4.The above point explains why Napoleon III used aluminum dinnerware for his

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~~Synthesis of Alum from Aluminum~~
stuff is my mid term grade alum
synthesis lab answers the synthesis of
alum lab michaela tonsager and kaili
johnson conclusion we determined
that our sample was in fact alum our
melting point of 994 degrees c was
similar to the published melting point

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of 925 degrees c our percent sulfate
was 4244

~~Analysis Of Alum Lab Answers Sulfate~~
SYNTHESIS OF ALUM Purpose: In this
lab you are going to carry out a series
of reactions to transform a piece of
aluminum foil into white crystals

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(alum). You will use stoichiometry to figure out the limiting reagents and yield for your reaction. Background: Aluminum is the third most common element in the earth ' s crust, but it forms very

~~SYNTHESIS OF ALUM~~

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Alum Synthesis Lab Answers The
Synthesis of Alum Lab Michaela
Tonsager and Kaili Johnson

Conclusion We determined that our
sample was in fact alum. Our melting
point of 99.4 degrees C was similar to
the published melting point of 92.5
degrees C. Our percent sulfate was

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42.44%, which is close to the The
Synthesis of Alum Lab by Michaela
Tonsager Lab #4 Synthesis of Alum
Purpose-Synthesize a type of alum

~~Alum Synthesis Lab Answers—
securityseek.com~~

This is a video I made during my last

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lab. Might get used in a lab report, not sure. The editing in this video is very basic. I just switched from Sony Movie...

This clearly written, class-tested

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manual has long given students hands-on experience covering all the essential topics in general chemistry. Stand alone experiments provide all the background introduction necessary to work with any general chemistry text. This revised edition offers new experiments and expanded

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information on applications to real
world situations.

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Popular Mechanics inspires, instructs and influences readers to help them master the modern world. Whether it ' s practical DIY home-improvement tips, gadgets and digital technology, information on the newest cars or the latest breakthroughs in science -- PM

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is the ultimate guide to our high-tech lifestyle.

BANNED: The Golden Book of Chemistry Experiments was a children's chemistry book written in the 1960s by Robert Brent and illustrated by Harry Lazarus, showing

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how to set up your own home laboratory and conduct over 200 experiments. The book is controversial, as many of the experiments contained in the book are now considered too dangerous for the general public. There are apparently only 126 copies of this book in

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libraries worldwide. Despite this, its known as one of the best DIY chemistry books every published. The book was a source of inspiration to David Hahn, nicknamed "the Radioactive Boy Scout" by the media, who tried to collect a sample of every chemical element and also built a

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model nuclear reactor (nuclear reactions however are not covered in this book), which led to the involvement of the authorities. On the other hand, it has also been the inspiration for many children who went on to get advanced degrees and productive chemical careers in

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industry or academia.

Popular Mechanics inspires, instructs
and influences readers to help them

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master the modern world. Whether it ' s practical DIY home-improvement tips, gadgets and digital technology, information on the newest cars or the latest breakthroughs in science -- PM is the ultimate guide to our high-tech lifestyle.

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