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 $\text{pH} = \text{pK}_a + \log \left[ \frac{\text{mmol Base}}{\text{mmol Acid}} \right]$   
 $\text{pH} = \text{pK}_a + \log \left[ \frac{0.65 \text{ mmol}}{0.65 \text{ mmol}} \right]$   
 $\text{pH} = \text{pK}_a + \log (1)$   
Therefore, when the weak acid is 50% neutralized,  $\text{pH} = \text{pK}_a$ .  
Step 3: Solve for the pH at the equivalence point.

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