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This excellent new text dispels the fear that the word electrochemistry commonly instils in chemistry students. Throughout the mathematical content has been left to a minimum for clarity, whilst retaining the important necessary physical insight.

The renowned Oxford Chemistry Primers series, which provides focused introductions to a range of important topics in chemistry, has been refreshed and updated to suit the needs of today's students, lecturers, and postgraduate researchers. The rigorous, yet accessible, treatment of each subjectarea is ideal for those wanting a primer in a given topic to prepare them for more advanced study or research. Moreover, cutting-edge examples and applications throughout the texts show the relevance of the chemistry being described to current research and industry. The learning features provided, including questions at the end of every chapter and online multiple-choice questions, encourage active learning and promote understanding. Furthermore, frequent diagrams, margin notes, further reading, and glossary definitions all help to enhance a student'sunderstanding of these essential areas of chemistry. This brand new addition to the series provides the most accessible first introduction to electrochemistry, combining explanation of the fundamental concepts with practical examples of how they are applied in a range of real-world situations.

This textbook provides a framework of the key concepts involved in electrochemical kinetics. A wide range of modern electrochemical techniques and applications are discussed. The mathematical content has been minimized for clarity, while retaining the important results necessary for physical insight. A substantial series of examples and illustrations is taken from the recent research literature to explore the potential applications of electrochemical techniques.

Another winning primer! This new addition to the popular series provides a basic introduction to equilibrium electrochemistry, focusing on electrode potentials and their applications. It builds on a knowledge of elementary thermodynamics giving the student an appreciation of the origin of electrode potentials and shows how these are used to deduce a wealth of chemically important information and data such as equilibrium constants, the free energy, enthalpy and entropy changes of chemical reactions, activity coefficients, the selective sensing of ions. It is mathematically simple, the emphasis throughout is on understanding the foundations of the subject and how it may be used to study problems of chemical interest.

Comprehensive text and reference covers all phenomena involving light in semiconductors, emphasizing modern applications in semiconductor lasers, electroluminescence, photodetectors, photoconductors, photoemitters, polarization effects, absorption spectroscopy, more. Numerous problems. 339 illustrations.

Electron transfer reactions are of fundamental significance in many areas of inorganic, organic and biological chemistry, and electrochemical techniques are a useful tool for studying them. This book provides an overview of recent advances in voltammetry and electrochemistry, broadening the scope of their application and suggesting new problems that they may be able to address in the 21st century.

Solvents other than water are used in chemical analysis, chemical manufacturing, and in specialized syntheses. This book covers the principles and uses of non-aqueous solvents at a level suitable for first or second-year undergraduates. The book first discusses the general properties of solvents, and introduces the necessary concepts for making rational choices of solvents for different applications. There is a discussion of the various chemical interactions between solvents and the substances dissolved in them, and how solvents change the course of reactions. The chemistry of 16 common solvents is discussed, emphasizing the advantages and disadvantages of each. The book concludes with an account of the chemistry of molten salts and discusses the use of low melting temperature compounds as synthetic media. The book expands on the brief treatment of non-aqueous solvents given in many textbooks while avoiding the complexities introduced in research treatises. It is the only book currently available that provides an in-depth treatment accessible to undergraduates.

This book will give students a thorough grounding in pH and associated equilibria, material absolutely fundamental to the understanding of many aspects of chemistry. It is, in addition, a fresh and modern approach to a topic all too often taught in an out-moded way. This book uses new theoretical developments which have led to more generalized approaches to equilibrium problems; these approaches are often simpler than the approximations which they replace. Acid-base problems are readily addressed in terms of the proton condition, a convenient amalgam of the mass and charge constraints of the chemical system considered. The graphical approach of Bjerrum, Hagg, and Sillen is used to illustrate the orders of magnitude of the concentrations of the various species involved in chemical equilibria. Based on these concentrations, the proton condition can usually be simplified, often leading directly to the value of the pH. In the description of acid-base titrations a general master equation is developed. It provides a continuous and complete description of the entire titration curve, which can then be used for computer-based comparison with experimental data. Graphical estimates of the steepness of titration curves are also developed, from which the practicality of a given titration can be anticipated. Activity effects are described in detail, including their effect on titration curves. The discussion emphasizes the distinction between equilibrium constants and electrometric pH measurements, which are subject to activity corrections, and balance equations and spectroscopic pH measurements, which are not. Finally, an entire chapter is devoted to what the pH meter measures, and to the experimental and theoretical uncertainties involved.

This is an introduction to the areas of application of electroanalysis, which has an important role with current environmental concerns, both in the laboratory and in the field.

This beginner's guide to cyclic voltammetry is designed to take you from novice to competent in a week. It bypasses all the mathematical proofs that often act as barriers to learning and begins with the practical information about experimental setup which will let you immediately start collecting and interpreting cyclic voltammograms. After the knowledge needed for gaining hands-on experience has been laid out, the underlying concepts that explain what happens at a molecular level during a cyclic voltammogram are described using easily understandable pictures and animations. This book is not meant to replace any of the go-to textbooks for electrochemistry, but to serve as a stepping stone on ones journey into the field, like a helpful postdoc in book form.